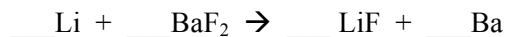


## Mass-Mass Stoichiometry

Write out the flowchart mass to mass stoichiometry that you have in your notes in the space below.

Balance the following reactions and then use them to answer the questions below.

**Show all your work with units.**



If 3.5 **grams** of Li are used in the above reaction, how many **grams** of LiF are produced?

1. What type of stoichiometry is this? \_\_\_\_\_
2. The first step is to convert grams of Li to mol of Li, do that below.
3. Now use a molar ratio from the balanced equation to convert from mol of Li to mol of LiF.
4. Finally, convert moles of LiF to grams of LiF using it's molar mass.

Follow the same procedures as 1-4 and use the flow chart above to answer the following questions:

5. In the above reaction, how many g of Ba are produced when 131 g of  $\text{BaF}_2$  are used?
6. If 35.0 grams of Li react, how many g of Ba are produced?